

Shortened version of the title: OXALATES ON MONUMENTS IN VENICE

OXALATE PATINAS ON STONE MONUMENTS IN VENETIAN LAGOON: CHARACTERISATION AND ORIGIN

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Abstract

Calcium oxalate patinas have been sampled on architectural elements of two ancient churches located on Torcello Island (Venetian Lagoon) and subsequently analysed. The site had been selected presenting patinas of exceptional amount and thickness, elsewhere generally found as thin alteration layers covering stone surfaces or within black crusts on monuments in urban environments. Optical and mineralogical analyses suggest that these patinas are not the result of a simple deposition process, but originate as surface “transformations” of the substrate, and are mainly composed of dihydrate calcium oxalate (weddelite) and gypsum. Among the experimental techniques, isotope analyses (C and S) have been specifically carried out aiming at achieving a better understanding of their origin and possible causes of formation. The observed carbon isotopic fingerprint reveals in fact a strong biological fractionation recorded in the oxalate patina ($\delta^{13}\text{C}$ ranging from -22.3‰ to -28.0‰), almost exclusively attributable to C3 pathway photosynthesis processes. Slightly negative sulphur isotope ratio values seem to be inversely correlated with patina exposure to atmospheric agents and pollutants. The results obtained prove

that calcium oxalate originates from a biomineralization process induced by lichen colonization in specific environmental conditions.

KEYWORDS: oxalate patina, cultural heritage, isotope analyses, biomineralization, lichen

1. INTRODUCTION

Stones of monuments and historic buildings experience different chemical, physical and biological processes during their exposure to the atmosphere which lead to the formation of a surface layer, depending on both the stone characteristics and the microclimate interacting with materials. A number of surface typologies have been identified in the literature on carbonate and silicate stones of the built cultural heritage in different geographical areas, and their origin has been linked to the effects of air pollution, rain leaching, biodeterioration, tourism pressure or conservation treatments (Viles et al., 1997).

Although investigation has been made concerning the composition and formation mechanisms of most damage layers found on historic stone monuments and buildings in polluted and unpolluted areas, an open problem still exists with regard to the origin of oxalate patinas.

Oxalate patinas have been found on cultural heritage from a variety of historical periods (from Greek-Roman to modern architecture) and on different materials (Del Monte et al., 1987). A comprehensive review and discussion on oxalate patina composition in diverse settings is presented by Matteini and Moles (1986) and Brimblecombe (1992), but most of the debate still focuses on stone monuments.

Calcium oxalate patinas are alteration layers characterised by variable colour, thickness and texture. Two main typologies are reported in the literature: thin layers, orange-brownish in colour, with high consistency and coherency, and white patinas of considerable thickness and scarce consistency (Ariño et al., 1995). Some patinas show a homogeneous structure, others a stratified one, but the mineralogical composition is in both cases very similar: monohydrate and dihydrate calcium oxalates (i.e. whewellite and weddellite) are the main crystalline phases associated with gypsum and quartz in different amounts and ratios.

The low solubility of whewellite and weddellite protects the underlying substratum against atmospheric weathering (Alessandrini, 1989). Therefore, it may be argued that the formation of calcium oxalate patinas mostly produces an aesthetic impact with chromatic alterations of the monument surface on which it develops.

As widely reported in the literature, three are the possible origins of calcium oxalate patinas on architectural surfaces in marble and limestone: biodeterioration, air pollution impact, degradation of organic treatments (Sabbioni et al., 2003; Vazquez-Calvo et al., 2007).

A considerable number of papers ascribe oxalate patinas to a biological origin. The chemical-physical interaction between colonizing microorganisms secreting oxalic acid and the limestone substrate leads to the dissolution of carbonates. The calcium cations freed by the process are involved into the formation of calcium oxalate (Del Monte and Sabbioni, 1983). This occurrence is confirmed by laboratory experiments which demonstrated that the interaction between oxalic acid (Kouzeli et al., 1996) or colonial bacteria (Monte, 2003) and samples of Carrara marble, produces calcium oxalate layers. The neoformation of thin oxalate layers occurs along the lichen-

substrate interface, where there are suitable pH, humidity and reduction potential values (Adamo et al., 1997).

In high humidity and basic pH conditions weddellite crystallizes preferentially, whereas in acid and hydrated environments whewellite is more stable (Frey-Wyssling, 1981; Ascaso et al., 1982; Horner et al., 1985).

Because of its low solubility, the calcium oxalate tends to amass in the thallus of the lichen to form a thin crystalline layer around it (Jackson, 1981), but also at the lichen-substrate interface (Ascaso and Galvan, 1976). Del Monte et al. (1987) reported that the interface which subdivides the lichens from the unaltered substrate is extremely irregular, implying a pervasive interaction between oxalate and carbonate (Ariño et al., 1995). Therefore, oxalate patinas cannot be considered as a simple “deposition” on the stone, but a “transformation” of the stone itself.

The literature also mentions weddellite and whewellite on basalt stone in natural outcrops as a consequence of lichen colonization (i.e. *Pertusaria corallina*). The action of mycobionts on primary minerals, such as Ca-rich plagioclases, allows the availability of calcium ions and the subsequent formation of calcium oxalate (Jones et al., 1980; Arocena et al., 2007). Lichens colonizing serpentinites lead to the precipitation of magnesium oxalate, as glushinskite (Wilson et al., 1980; Wilson et al., 1981; Adamo et al., 1993; Favero-Longo et al., 2005). The neoformation of Mg-oxalate has also been encountered during *in vitro* experiments on fungi-mineral interaction on dolomite (Kolo and Claeys, 2005). Lichen colonization on carbonate structures produces not only calcium oxalates, but also metallic oxides and hydroxides, as well as aluminium-silica compounds, may form (Adamo and Violante, 2000).

De Santis and Allegrini described, in 1989, oxalic acid associated to air pollution deposition as a second hypothesis for oxalate patina origin. Sabbioni et al. (2003) focused a review on small (C1-C2) organic anions, including oxalate, in black crusts on monuments and buildings. They showed that these short chain ions are almost constantly associated with sulphur and nitrogen compounds linked to airborne gas and particle deposition emitted by combustion sources and producing black crust formation on stone substrates. In aerosol composition the positive correlation between non sea salt sulphate and oxalate could indeed be ascribed to similar pathway during in-cloud formation (Yu et al., 2005 and reference therein).

As a third hypothesis, the decomposition of organic treatments applied in the past to protect the surface of monuments is considered to be an origin of calcium oxalate formation. Since the classical antiquity, animal- and vegetal-derived components have been used as protective compounds (Rampazzi et al., 2004; Gabrielli and Parenti, 1992; Cennini, 1991; Frazzoni, 1988; Piva, 1988; Fassina, 1989). Berlucchi and Ginanni Corradini (1996) interpreted the presence of joints in some oxalate patinas, recognized by thin section optical microscope analyses, as post-application pull back structures of the protective treatment layer. Moreover, infrared spectrometer analyses (FT-IR) of the same samples revealed the presence of barium silicates used in the past to whiten the compound.

Oxalate patinas have also been found indoors, on surfaces not directly exposed to atmospheric weathering. The few studied cases ascribe the presence of calcium oxalate mainly to the alteration of protective treatments (Alessandrini, 1989; Chiari et al., 1996).

The identification of the origin of oxalate patinas on stone monuments continues to be a challenging area of research, in spite of more than twenty years of debate. Today, the growth mechanism of oxalate patinas on monuments may be gaining in importance, in view of the different future trends assumed in the near and far future by climate parameters linked to biodeterioration (i.e. temperature and precipitation), pollution and anthropic pressure (Grossi et al., 2008; Bonazza et al., 2009a; Bonazza et al. 2009b, Gomez-Bolea et al., 2012). Within the perspective of global change, the need for an effective analytical technique with which to solve the remaining problem of oxalate patina formation mechanisms is therefore particularly pressing in the field of cultural heritage conservation.

Stable carbon isotopic ratios due to the precise fingerprint drawn by biological processes in both polluted and unpolluted areas may be a suitable indicator for interpreting the formation process of these specific surface layers. With this perspective, the aim of the present work is, besides an exhaustive chemical and mineralogical characterization of the alteration patinas onto monuments in Torcello Island, to explore if isotope analyses can constitute a powerful tool for discriminating the origin of surface oxalate patinas formed on heritage building stones.

The work focuses on the results achieved by a first-time application of isotope analyses to this particular typology of neoformation surface layer, with the aim of providing a contribution to the identification of oxalate patina genesis.

2. MATERIALS AND METHODS

2.1. Sampling

The sampling was performed on Torcello island, located in the Venetian Lagoon, 9 km north-east of Venice (Figure 1). The site was selected because of the outstanding abundance and thickness of calcium oxalate layers on the surfaces of the island's monuments, allowing both a complete characterization of this specific typology of surface layer and the application of isotope analysis, which requires homogeneous surface layers. The samples were collected from the columns and capitals of different lithologies belonging to the external perimeter loggias of the two churches of Santa Maria Assunta (Figure 2a) and Santa Fosca (Figures 3a and b). Most of the columns are constituted by impure greenish marble (Greek Marble s.s.) and nodular reddish limestone (*Rosso di Verona*, Figure 2b), only two being composed of micritic limestone. The description and localization of the samples on which the experimental work was based are summarized in Table 1.

Most of the patinas observed on the columns and capitals of the two churches were white-grey in colour, with a thickness of up to 0.5 cm, scarcely coherent and revealing a plastic behaviour during sampling (Figure 4).

2.2. Analytic procedures

Mineralogical and petrographic investigation of the samples was carried out by an Olympus BX51 polarised light optical microscope (OM) and X-ray diffraction (XRD) using a Philips

PW1710 Based diffractometer, equipped with Cu pipe ($K\alpha_1 = 1.54060 \text{ \AA}$, $K\alpha_2 = 1.54439 \text{ \AA}$). Morphological observations were performed by a Philips XL 20 scanning electron microscope (SEM), equipped by x-ray energy dispersion analyzer (EDX) to perform elemental analyses. Differential and gravimetric thermal analyses were carried out by a METTLER TOLEDO TGA/SDTA 851, equipped with TSO 800GC1 programmable gas switch. A combination of these techniques was employed to attain an accurate characterization of the surface layer and to test its homogeneous composition.

With the aim of quantifying the anion concentrations in the collected samples, ion chromatographic analyses were performed using a Dionex Chromatograph (model 4500i), equipped with conductivity detector (Dionex CDII), following the methodology reported by Gobbi et al. (1995).

Isotopic analyses were carried out by a VARIO EL element analyser coupled with an IR MS mass spectrometer CNS (accuracy of 0.2 ‰). The samples were powdered in an agate mortar and then treated with 37% HCl solution in order to eliminate carbonate phases, thus avoiding the carbon signature of the substrate (marble and limestone). The CO_2 freed by carbonate dissolution were trapped by NaOH grains within the stagnate glass dryer where the samples were stored.

3. RESULTS AND DISCUSSION

3.1. OM analysis

Thin sections of oxalate patinas, collected from different lithologies, and one of substrate (a marble column fragment) were analysed by optical microscope. In this way, two kinds of oxalate patina were recognised, in relation to the different substrate on which they form. The first, collected from architectural elements in limestone (*Rosso di Verona*), is mainly composed of a yellow-brownish microcrystalline matrix of calcium oxalate that included rare uncoloured phyllosilicate (muscovite) and small rounded black oxides (Figures 5a and b). The measured average thickness is around 1100 μm . The second kind, collected from impure marble columns and capitals, shows a similar microcrystalline matrix to the former, and includes a remarkable amount of phyllosilicate, calcite and quartz crystals (Figures 6a and b). In the figure, it is possible to identify a patina's growth pattern perpendicular to the surface layer-substrate interface, reflected by the isorientation of the phyllosilicate crystals. The average thickness measured in this case is around 900-950 μm .

The substrate sample shows an inequigranular texture with isorientation of the mineral grains typical of metamorphic processes. It was found to be mainly composed of euhedral white calcite, under plane polarized light with the characteristic rhombohedral cleavage, and of greenish elongated chlorite crystals. These fibrous minerals are of discrete dimensions, and are homogeneously present in the section, showing low interference colours by cross polarized light, typical of Mg-rich chlorites (pennine). Quartz is also present in a smaller amount in anhedral

crystals with undulated extinction. Elongated muscovite crystals constitutes the second phyllosilicate phase recognised. On the basis of the results obtained, the substrate sample was classified as low grade metamorphic rock (impure marble). Following the petrographical characterization, it may be argued that the significant presence of phyllosilicate, calcite and quartz found on the oxalate patina collected from the column in impure marble of Santa Maria Assunta Cathedral is linked to the underlying rock. Particularly in the process of alteration.

3.2. SEM-EDX analysis

Scanning electron microscope (SEM) analysis were performed on three samples of oxalate patinas (TORNEW5, TORNEW6, TORNEW7), in order to investigate their morphology from micro to nano-scale and, through the x-ray energy dispersion (EDX) probe, to evaluate their elemental composition.

In all the investigated samples, the patina's surface was mostly composed of a homogeneous microcrystalline matrix, with globular morphology and diffuse porosity, sometimes broken by fractures. The EDX analyses of the surfaces showed them to be mainly composed of Ca, with smaller amounts of Si, Cl, S, and Al, while no particles from atmospheric deposition were found to be present (Figures 7a and b).

Increasing magnification led to the identification of phyllosilicates, including muscovite (Figures 8a and b) and feldspar crystals of about 100 μm length in samples TORNEW5 and TORNEW6, where smaller amphiboles, chlorite and cubic halite were also found. As already evidenced by OM observation the identified phyllosilicate are likely to be related to the accessory minerals of

the substrate (impure marble) more than to atmospheric soil dust deposition. This finding allows us to advance the hypothesis that in the oxalate formation process, silicate minerals belonging to the substrate, such as micas and quartz, are embedded in the newly formed patina without undergoing any substantial chemical alteration. In addition in the analysed samples it is worth noting the presence of organic filaments emerging by about 100 μm from the surface. A large number of smaller fragments very similar to fungal hyphae appears networked within the matrix (Figures 9a). In sample TORNEW7, these organic fragments are covered by a thin layer of calcium oxalate (Figure 9b).

3.3. XRD analysis

The X-ray diffraction analysis of five samples of calcium oxalate patina was carried out, in order to determine their mineralogical composition by means of a qualitative and semi-quantitative approach. The obtained results are reported in Table 2.

It should be noted in the table, that most of the patinas are mainly composed of tetragonal calcium oxalate dihydrate and subordinate gypsum. Minor amounts of quartz and calcite are usually recognized, whereas halite, iron and manganese oxide/hydroxide are always present in traces. As well as calcium oxalate above discussed, the formation of gypsum in these patinas can be ascribed to several origins. Dry atmospheric deposition of SO_2 on carbonate substrates followed by rainwater acidification processes lead to the formation of $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ due to the dissolution of calcite that releases calcium cation. Moreover, it should be pointed that aerosol sulphate concentrations have been found in positive correlation with oxalate in different environments, suggesting the possibility of a common in-cloud pathway for these compounds

(Crahan et al., 2004; Yu et al., 2005 and reference therein), and pointing to a common atmospheric source of the gypsum and calcium oxalate in patinas. Sea spray deposition of sulphate, usually occurring in brackish environments, is another possible source in the formation of this compound. Phyllosilicate (chlorite, mica, illite, clinocllore) were found to be present in some of the samples (TOR4bis, TOR5) collected from marble columns.

The constant presence of halite is due to sea spray recrystallization, whereas iron, magnesium and manganese oxide/hydroxide could derive from both the substrate and deposited atmospheric aerosol. However, a number of authors interpret these phases as neoformation minerals linked to the interaction between lichen biological activity and rock substrate through the mobilization of metallic ion (Barker and Banfield, 1996; Prieto et al., 1997; Adamo and Violante, 2000; Arocena et al., 2003). In particular, the presence of wustite in one sample (TOR11) should be ascribed to “de-dolomitization” processes, leading to the oxidation of bivalent iron contained in the substrate.

Calcium oxalate monohydrate appeared only in sample TOR2bis, probably due to crystallization of calcium oxalate in both low humidity environments or low pH conditions (Horner et al., 1985). In this case, subsequent hydration, which leads to the transformation into the dihydrate form, could therefore not be completed.

3.4. DTA-TGA analysis

Differential thermal coupled with gravimetric analysis was conducted on two samples (TOR4bis and TOR8) representative of the two oxalate patina types. The analyses were performed under

O₂ constant flux conditions. In matrixes containing calcium oxalates, the quantitative interpretation of DTA-TGA results is particularly complex, because hydrated oxalate forms (mono, bi- and trihydrates) loose water crystallization at about 200 °C, while on reaching 500 °C they decompose into CaCO₃, developing CO₂, as highlighted by Riontino et al. (1998). The data obtained showed the presence of organic carbon with a exothermic peak at 310 °C, producing a weight loss of between 11 and 14 wt%. The decomposition of anhydrous oxalate (CaC₂O₄) into calcite (CaCO₃) occurs through a exothermic reaction at 460 °C, with the loss of 9 wt% in TOR4bis and 14 wt% in TOR8. Decomposition of carbonate phases at temperatures between 760-780 °C is coupled with loss of weight of 14% (TOR4bis) and 16% (TOR8). It should be emphasised that the weight loss incurred by carbonate phase decomposition give an overestimation of calcite, because of the above mentioned calcium oxalate decomposition at 460 °C.

3.5. IC analysis

The IC analyses provided the concentration of the ion soluble fraction of the patinas. The results obtained indicate that oxalate anion concentration ranges from 1 to 7%. The experimental data proved that the sampled patinas are formed by considerable higher amounts of oxalates than those measured in damage layers from monuments in different locations, built in different materials, which were constantly lower than 1.6 % (Riontino et al., 1998; Sabbioni et al., 2003), with exception of black crusts sampled in Venice, which showed a maximum concentration of 3.4 % (Bonazza et al., 2005). These results demonstrate that the selected samples were particularly suitable for testing isotope analyses.

In addition to oxalate, other organic anions were found: acetate and formate with concentrations of one and two order of magnitude lower than oxalate anion, being respectively 0.4% and 0.001%.

Among the non-organic anions, sulphate predominates followed by nitrate and chloride. Phosphates are present in variable concentrations, and nitrate and fluoride varying from trace to negligible concentrations.

3.6. Isotopic analysis

The isotope analyses were performed on three surface patina samples from the church of Santa Maria Assunta, collected from impure marble (TORNEW5 and TORNEW6) and limestone (TORNEW7), both internal and external. The data obtained for $\delta^{13}\text{C}$ and for $\delta^{34}\text{S}$ are reported in Table 3.

The values of $\delta^{13}\text{C}$ ranging from -22‰ to -28‰ indicate that carbon fractionation is not affected by different substrate lithology or the position of the patina. Considering that the $\delta^{13}\text{C}$ in carbonate rocks varies from -1 to 6‰ , and that the only carbon source able to affect this value is atmospheric CO_2 with average $\delta^{13}\text{C}$ equal to -7‰ , it can be concluded that the only process in nature responsible for the large negative $\delta^{13}\text{C}$ recorded on the Torcello oxalate patinas, is biological carbon fractionation. Colonising organisms, i.e. lichens and fungi residing on rocky substrates, produce similar paragenesis of mineral neoformation having different carbon isotopic fractionation. Kolo et al. (2007) observed slightly negative $\delta^{13}\text{C}$ (-4.2‰) in Mg and Ca oxalates neomineral originating from fungi attack on limestone and dolomite substrates ($\delta^{13}\text{C} \approx +2.1\text{‰}$).

ca.). More negative $\delta^{13}\text{C}$ values have been reported for lichen biological activity. Using atmospheric carbon dioxide as the only nutrient for the photosynthesis process (C3 pathway), the lichen phycobiont (green or blue algae) preferentially select $^{12}\text{CO}_2$ instead of $^{13}\text{CO}_2$ through a diffusion process. Such fractionation results in $\delta^{13}\text{C}$ around -27‰ for higher plants adopting this kind of photosynthesis cycle (Fogel and Cifuentes, 1993), whereas this value can range between -17‰ and -34‰ for lichens depending on the intensity of the diffusion process (Lange and Ziegler, 1986; Hoefs, 1997). In addition, lichens living in more moist conditions are depleted in ^{13}C compared to those growing in relatively drier environments (Russ et al., 1996; Brugnoli et al., 1988; Beazley et al., 2002 and references therein). Beazley et al. (2002) found that calcium oxalate pruina is enriched in ^{13}C , compared to lichen tissues producing higher mean values of $\delta^{13}\text{C}$ of between 6 and 7‰.

Based on such assumptions, it may be argued that TORNEW5, undergoing greater carbon fractionation, was sampled in the area mostly exposed to moist conditions: values in neomineral calcium oxalate were extremely low ($\delta^{13}\text{C} = -28\text{‰}$), implying that the lichen tissues potentially have mean values of $\delta^{13}\text{C}$ ranging from -34 to -37‰ , which are ascribed in the literature to the presence of photosymbiodemes, a more generalist lichen association (Maguas et al., 1993, 1995; Smith and Griffiths 1996). Assuming that atmospheric CO_2 is the only source of carbon utilised by lichen biological activity, excluding the uptake of substrate rock carbon (Beazley et al., 2002), the other possible cause of such low values of $\delta^{13}\text{C}$ is the starting value of atmospheric CO_2 . Widory and Javoy (2003 and reference therein) found an inverse correlation between the concentration of CO_2 in the atmosphere and its $\delta^{13}\text{C}$ value, asserting that all anthropogenic carbon dioxide contributions are ^{13}C depleted.

Sulphur isotope ratio $\delta^{34}\text{S}$ is slightly negative, ranging from $-0,2\text{‰}$ (TORNEW7) to $-2,5\text{‰}$ (TORNEW5). Sulphate of oceanic sea spray ($\delta^{34}\text{S} = +21\text{‰}$; Rees et al., 1978) or biogenic sulphur (DMS, $\delta^{34}\text{S} = +15 \pm 3,1\text{‰}$; Calhoun et al., 1991) are characterized by much heavier compositions. It has to be underlined that even if negligible, a sulphur contribution from sea salt deposition in the patinas cannot be completely excluded (aerosol marine effect is proved by halite determination by XRD and SEM analyses). Gypsum has already been recognized by some authors as biomineral, in association with calcium oxalate, related to the biological activity of lichens in unpolluted (Prieto et al., 1999) and polluted areas (Garty et al., 2002), but no isotopic data have been reported. Biological fractionation should not affect the sulphur isotope ratio, particularly in epiphytic lichens and mosses, where the values of $\delta^{34}\text{S}$ reflect the composition of atmospheric SO_2 , suggesting that no discrimination between ^{32}S relative to ^{34}S is involved during metabolic cycles (Krouse 1977; Winner et al. 1978; Case and Krouse 1980). The values recorded in the patinas, given the widespread lichen colonization on Torcello (Zitelli and Salvadori, 1982), led us to infer that the local effect of oil burning power plants in the vicinity should be limited to an atmospheric SO_2 depletion in ^{34}S , without a considerable increase in SO_2 concentration. Indeed, lichens living in excess of atmospheric SO_2 concentration proportionally enhance the production of volatile H_2S weighting the sulphur residue (Krouse, 1977; Krouse et al., 1991; Newman et al., 1991). The lowest $\delta^{34}\text{S}$ values found in lichens living in similar conditions on Newfoundland by Wadleigh (2003) are partially comparable (slightly higher e.g. $\delta^{34}\text{S} \approx 4\text{--}7\text{‰}$) to those encountered in the Torcello patinas, the main difference being the healthy condition of the lichens at the Italian site.

4. CONCLUSION

The analyses performed on the patinas from the loggias of two investigated churches on Torcello island reveal that they are mainly composed of dihydrate calcium oxalate (weddellite). The particular abundance of the metastable form of this mineral reflects the specific environmental characteristics of the studied site: a lagoon environment characterized by high relative humidity and sea spray deposition.

The integrated methodological approach carried out allowed us to assume a major biological origin of the calcium oxalate patina sampled. The SEM analyses performed in particular show a significant presence of fungal hyphae, suggesting a direct involvement of present or past biological activity in the formation of the patinas. The results led also to the identification of a thin microcrystalline weddellite/gypsum veil that sometimes covers the hyphae, pointing to a possible extracellular mineralization already observed in lichen thallus (Garty et al., 2002). The study of sample thermal behaviour confirms a significant presence of organic fraction and, in addition, ion chromatography shows a remarkably high concentration of oxalate anions.

Isotope analyses were performed to investigate the origin of surface oxalate layers on monuments. The isotope measurements for $\delta^{13}\text{C}$ of the oxalates in the patinas, indicate photosynthesis and a fixation processes based on the C3 pathway (Calvin-cycle), involving the Rubisco enzyme. Small variations can occur due to biogeochemical interactions on the microscale and different diffusion with related fractionation prior to fixation. In addition, $\delta^{34}\text{S}$ measurements suggest that a correlation between atmospheric exposure of patinas and sulphur

isotope ratio might exist. Even if it is impossible to exclude a marine contribution to patina formation, gypsum can also be associated with Ca-oxalates, as a biomineral product deriving from lichen activity in polluted environments. The possible marine contribution is limited by the isotopic fingerprint that should be characterized by heavier sulphur isotopic ratios. Isotopic results confirm that calcium oxalate, the main components of the Torcello island alteration patina, are related to a biomineralization process induced by colonizing lichens in a peculiar environment.

Therefore, the results obtained contribute to the debate on the origin of oxalate patina, demonstrating, for the first time using quantitative data, the biological origin of oxalate patina growth on stone monuments. In addition, the present work proves isotopic analyses to be a suitable methodology for discriminating the origin of oxalate patinas, which application should be further developed. Within a perspective of changing climate, the identification of the growth mechanisms of surface layers on monuments is essential for the purpose of developing suitable strategies for the protection of cultural heritage.

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Figure 1. Map of the Venetian lagoon illustrating the position of the sampling site of Torcello island, Venice, and the power plant and oil refinery of Porto Marghera.



Figure 2a. Perspective view of the church of Santa Maria Assunta (7th century).



Figure 2b. Santa Maria Assunta loggia before the church gate. Columns R1, R2, R6 are made of nodular limestone (*Rosso di Verona*), column R3 is made of impure marble (*Greek marble s.s.*), and column R4 and R5 are in limestone.

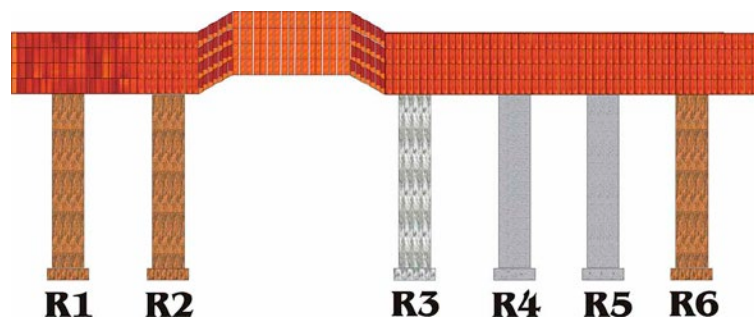


Figure 3a. Frontal view of Santa Fosca (11th century).



Figure 3b. Octagonal floor plan of Santa Fosca and of its loggia. The sampled columns are in marble.

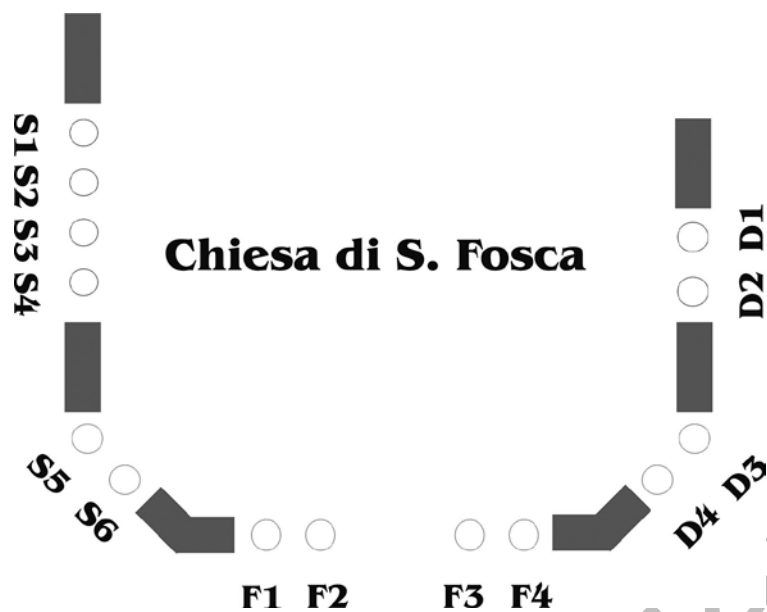


Figure 4. Sample of oxalate patina collected from a column in marble at Santa Fosca church.

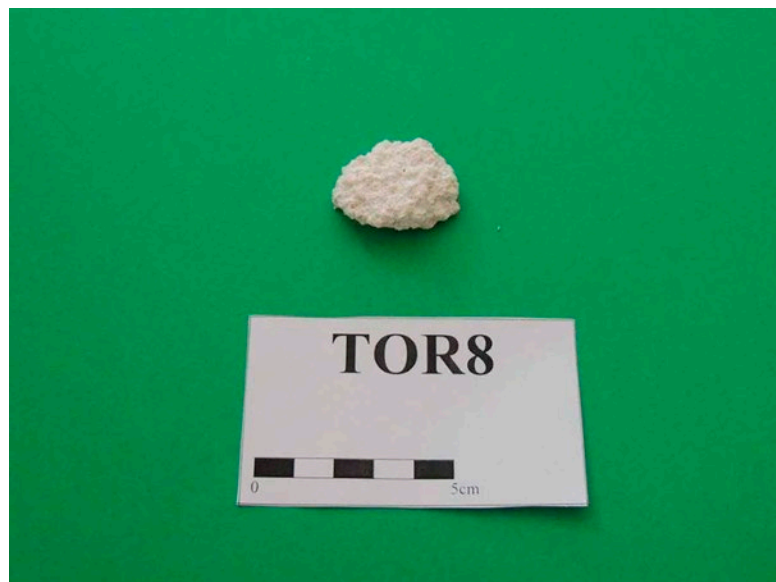


Figure 5a. Sample TOR2bis: microphotograph showing the high porosity of the oxalate patina (cross polarized light).



Figure 5b. Sample TOR2bis: small phyllosilicate crystal incorporated in homogeneous oxalate patina (cross polarized light).

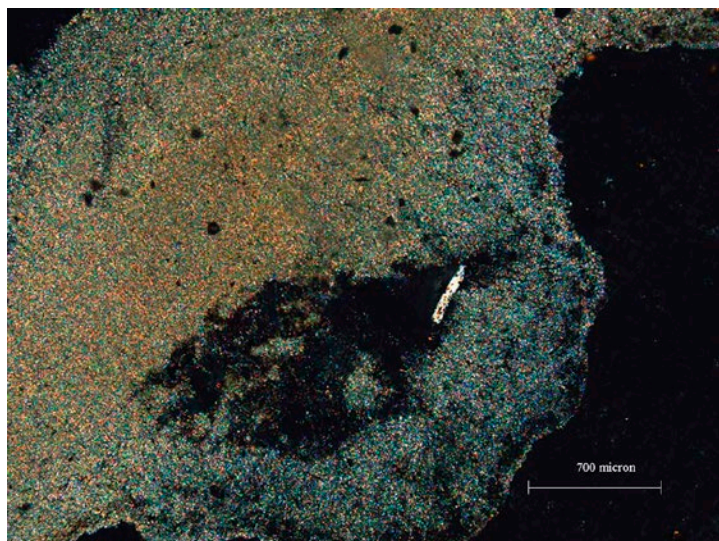


Figure 6a. Sample TOR4bis: microphotograph showing an orthogonal isorientation of phyllosilicates highlighting the growth direction of oxalate patina (cross polarized light).

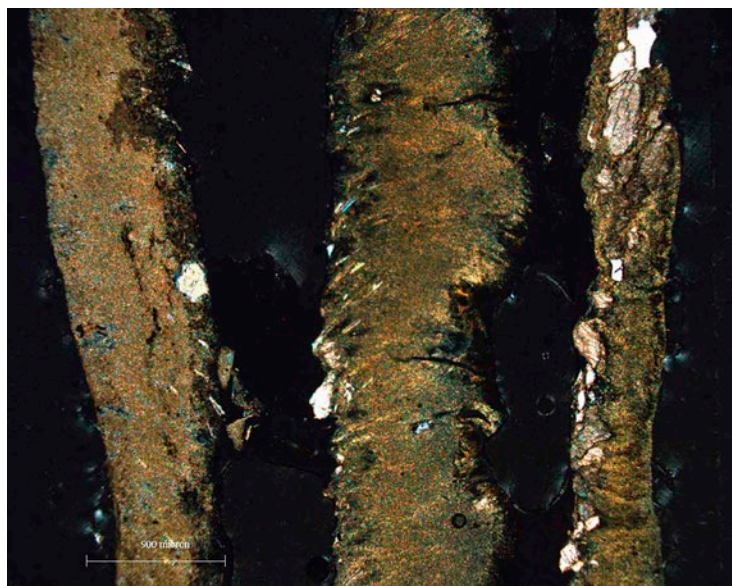


Figure 6b - Sample TOR4bis: microphotograph highlighting the presence of quartz crystals (white-yellow color) and calcite (pink color, on left) (cross polarized light).



Figure 7a. SEM microphotograph showing the patina morphology of sample TORNEW5.

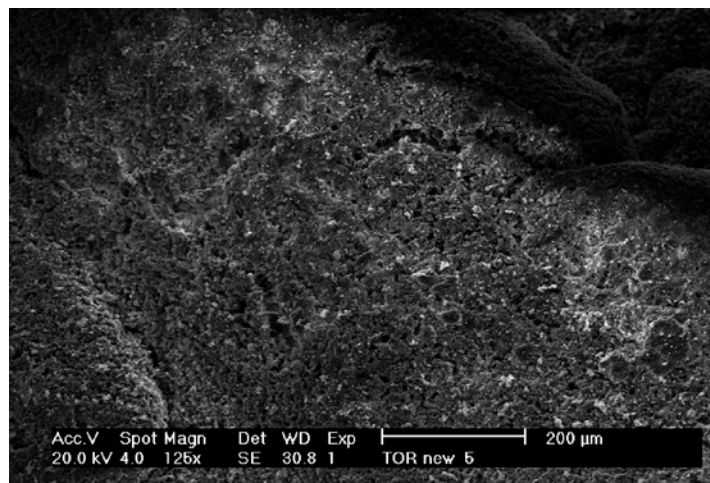


Figure 7b. EDX spectrum of TORNEW5 patina surface.

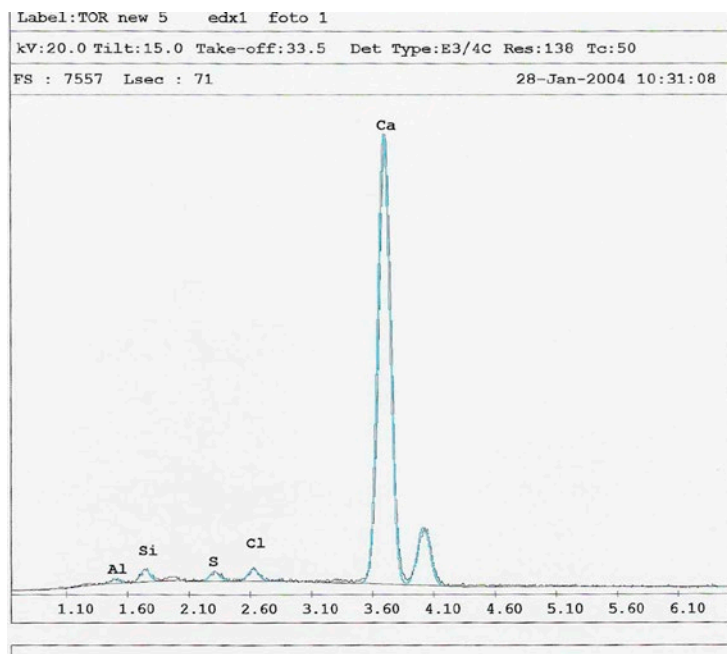


Figure 8a. SEM microphotograph of sample TORNEW5 showing a pair of muscovite crystals.



Figure 8b. EDX microanalysis of crystals in Figure 8a.

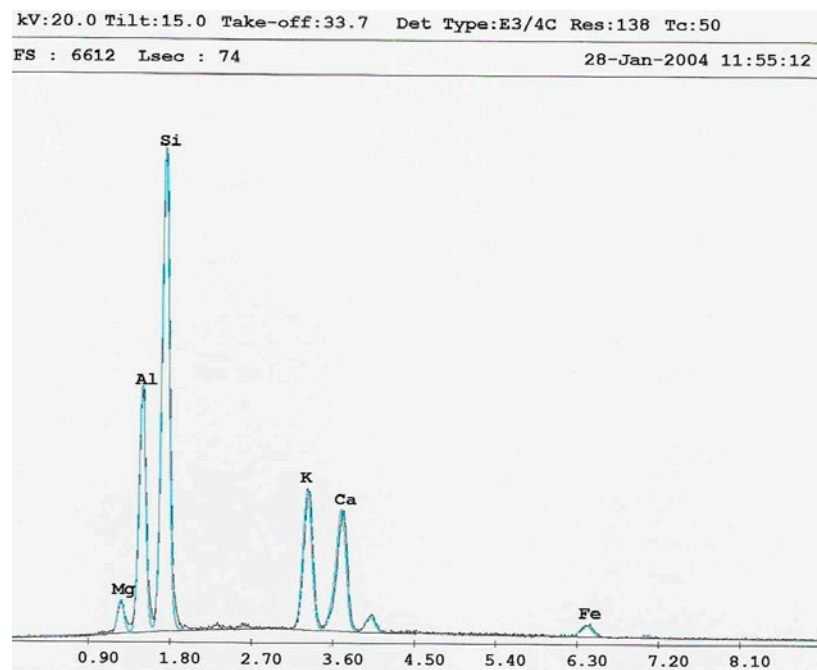


Figure 9a. SEM microphotograph of sample TORNEW5 showing numerous organic filaments resembling fungal hyphae with crystals on surface.

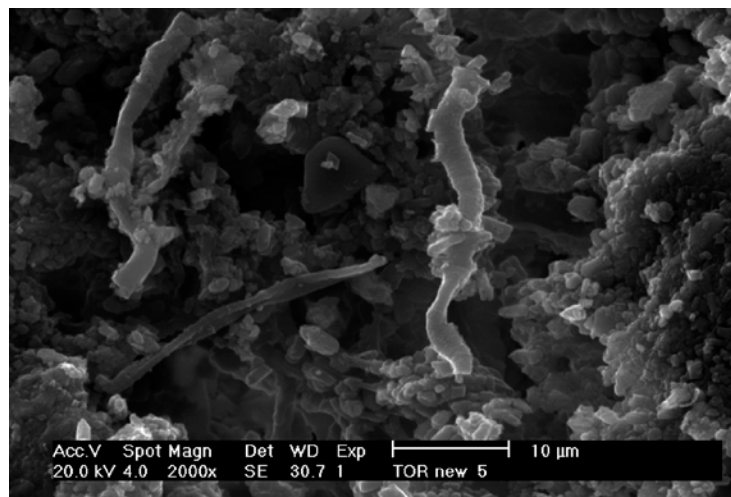


Figure 9b. SEM microphotograph of sample TORNEW7 showing numerous organic filaments resembling fungal hyphae covered by the layer of calcium oxalate.

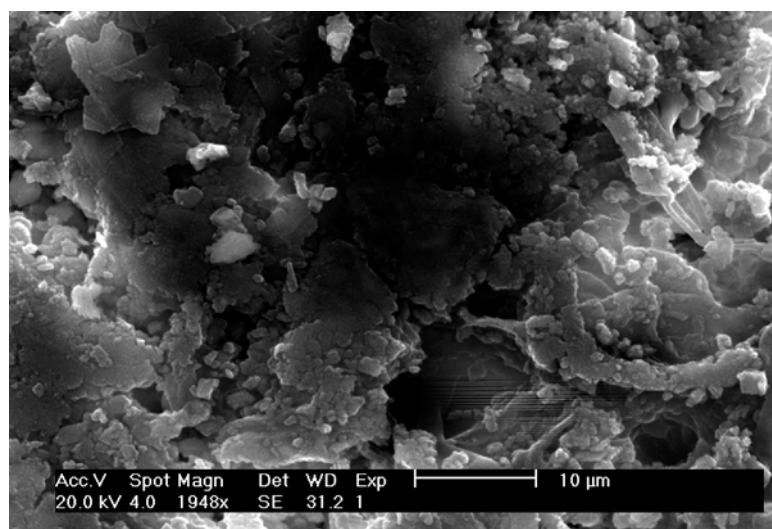


Table 1. Description and localization of samples from the two churches of Torcello Island (Venice)

Sample	Localization	Description
TOR2bis	Santa Maria Assunta Cathedral. Column R2. Collected from various points all over the surface	Pink (bottom side) - grey patina 2 mm thickness scarcely coherent
TOR3-TORNEW3	Santa Maria Assunta Cathedral. Column R3. Collected from various points all over the surface	White (bottom side) – light grey patina about 1 mm thickness scarcely coherent
TOR4bis-TORNEW4	Santa Maria Assunta Cathedral. Column R3. Collected only from inner side	White patina of about 2-3 mm thickness scarcely coherent
TOR5-TORNEW5	Santa Maria Assunta Cathedral. Column R3. Collected only from outer side	White (bottom side) – grey patina about 1 mm thickness, scarcely coherent
TORNEW 6	Santa Maria Assunta Cathedral. Column R3. Collected only from inner side	White (bottom side) – grey patina of about 2-3 mm thickness. Portion of conspicuously altered substrate was also collected during sampling
TOR7	Santa Maria Assunta Cathedral. Column R3	Marble fragment belonging to the column
TORNEW7	Santa Maria Assunta Cathedral. Column SMA1 in <i>Rosso di Verona</i> belonging to the inner loggia	Greyish patina of about 2 mm thickness scarcely coherent
TOR8	S. Fosca Church. Capital of column D2 in marble	White patina scarcely coherent
TOR11	S. Fosca Church. Upper part of Abside behind the church (marble)	White patina

Table 2. Semi-quantitative mineral phases estimated by x-ray diffraction (XRD) of surface patina samples (Wd=weddellite, Gy=gypsum, Q=quartz, C=calcite, Ch=chlorite, Mi/I=mica/illite, H=halite, Cc=clinoclone, Do=dolomite, Pe=periclase, Wu=wustite, Pc=pirocroite, Py/Mn=pirolusite/manganosite, Wh=whewellite)

Sample	Wd	Gy	Q	C	Ch	Mi/I	H	Cc	Do	Pe	W	Pc	Py/M	W
TOR2b	++++	++	±	±			±			±	±	±		±
TOR4b	++++	+++	++	++++		+	±	+		+	±	±	±	
TOR5	++++	++	+++	±	++	++	±			±	±			
TOR8	++++	+++	±	±			±				±	±	±	
TOR11	++++	+++		±			±		+	±	±	±	±	

Legend: +++++ = principal mineral present; +++++ = very much abundant; +++ = abundant; ++ = scarcely abundant; + = present; ± = traces

Table 3. Carbon and sulphur isotopic ratios and elemental abundance measured on three representative samples of calcium oxalate patinas from Torcello island

Sample	$\delta^{13}\text{C}$ (‰)	C (%)	$\delta^{34}\text{S}$ (‰)	S (%)
TORNEW5	-28.0 ± 0.2	17.0	-2.5 ± 0.2	1.9
TORNEW6	-22.3 ± 0.2	6.8	-1.1 ± 0.2	4.8
TORNEW7	-25.9 ± 0.2	19.5	-0.2 ± 0.2	1.8